

Postulates of the Kinetic-Molecular Theory

1. Particle Volume

- Volume of each gas particle *extremely small* compared with volume of container.
- Each gas particle has mass, but no volume.

2. Particle Motion

- Gas particles move in constant, random, straight-line motion until they collide with other gas particles, or container walls.

3. Particle Collisions

- All collisions are elastic – energy exchanged when gas particles collide with other gas particles, or with the container walls, but no energy lost through frictional forces.
- Between collisions, gas particles do not experience forces of attraction with each other, or with the container walls.
- Total kinetic energy of collection of gas particles is constant.