General Procedure for Writing Lewis Structures

1. Draw atoms (central atom + surrounding atoms) relative to each other.
   
a. For AB\(_n\) compounds, draw atom with *lower group number* in center (usually, this is the atom with the *lower* electronegativity).
   
b. If atoms have same group number, draw atom with *higher period number* in center.
   
c. H forms only one bond (i.e., never a central atom).

2. Sum the number of valence e\(^-\) of all atoms (= total number of valence e\(^-\) available).
   
a. For anion, *add* one valence e\(^-\) for each negative charge.
   
b. For cation, *subtract* one valence e\(^-\) for each positive charge.

3. Draw single (covalent) bond from each surrounding atom to the central atom.

4. For each single bond drawn, subtract 2 e\(^-\) from total number of valence e\(^-\) available.

5. Distribute remaining valence e\(^-\) in pairs so that each atom has 8 e\(^-\) (2 e\(^-\) for H).
   
a. First, place unshared e\(^-\) pairs on *surrounding atoms*.
   
b. Place any remaining e\(^-\) on *central atom*.

6. If central atom does not have 8 e\(^-\), change unshared e\(^-\) pair from one of surrounding atoms into bonding pair to central atom to make multiple bond.

7. Check that each atom has 8 valence e\(^-\) (2 e\(^-\) for H).