1. What is the oxidation number and charge of:
   \[
   \begin{align*}
   \text{Pb in PbO} & \quad +2 \\
   \text{C in CO}_2 & \quad +4 \\
   \end{align*}
   \]

2. In the following reaction
   \[
   \text{C + 2 PbO} \rightarrow \text{CO}_2 + 2 \text{ Pb}
   \]
   Identify the electron donor \( \text{C} \) (Give the reason for your answer)
   \[
   \text{Carbon goes from 0 to +4 charge}
   \]
   Identify the electron acceptor \( \text{Pb} \) (Give the reason for your answer)
   \[
   \text{Lead goes from +2 to 0 charge}
   \]
   Identify the oxidizing agent \( \text{Pb} \) (Give the reason for your answer)
   \[
   \text{Lead gained electrons}
   \]
   Identify the reducing agent \( \text{C} \) (Give the reason for your answer)
   \[
   \text{Carbon lost electrons}
   \]